

# Zero Administration Kit for Windows

## AN IMPLEMENTATION GUIDE

- How to reduce your administrative overhead
- Complete installation and configuration guidelines
- Tips for IT administrators in the trenches
- Lab exercises and troubleshooting tips

Michael McInerney



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# Zero Administration Kit For Windows

**Ensheng Dong**



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subsequent  $\Delta h$  values  
2  $\text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2}\text{O}_2(\text{g})$   $\Delta h = 227 \text{ kJ}$   
2  $\text{H}_2(\text{g}) + \text{C}_2\text{H}_6(\text{g}) \rightarrow \Delta h = 94.5 \text{ kJ}$   
 $\text{H}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \rightarrow \Delta h = 71.2 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2}\text{O}_2(\text{g}) \rightarrow 2\text{CO}_2(\text{g}) + 3\text{H}_2\text{O}(\text{g})$   $\Delta h = 283 \text{ kJ}$

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web thus for reaction 9 2 in which 1 mole of  $h_3o^+$  and 1 mole of  $oh^-$  combine to form 2 moles of  $h_2o$   $\Delta h = 55.8 \text{ kJ}$  as shown note that the enthalpy change  $\Delta h_{products} - \Delta h_{reactants}$  is positive if heat is absorbed that is if  $\Delta h_{products}$

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