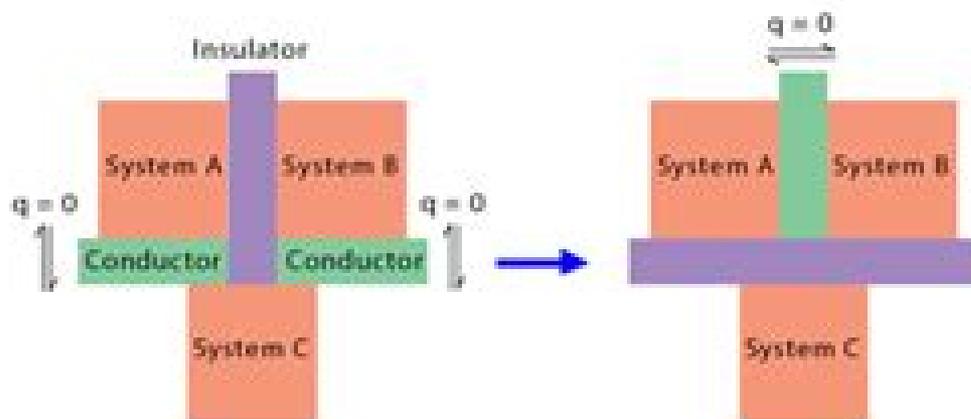


# Laws of Thermodynamics

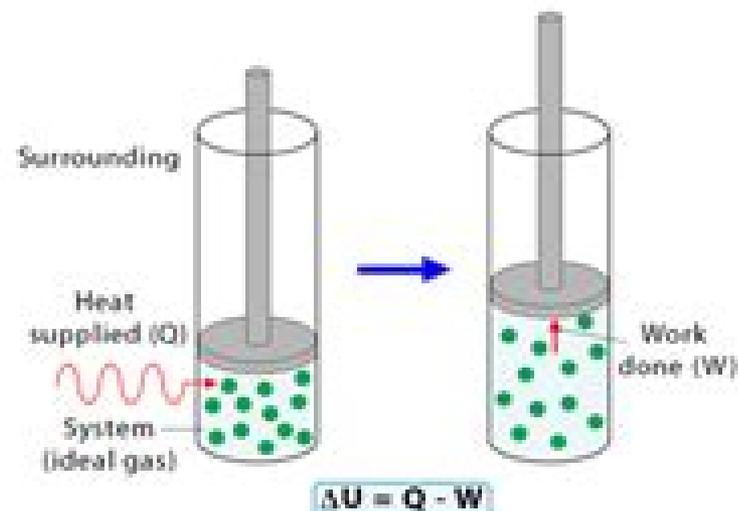
## Zeroth Law

If two thermodynamic systems are in equilibrium ( $q = 0$ ) with a third, then the two are in equilibrium with each other



## First Law

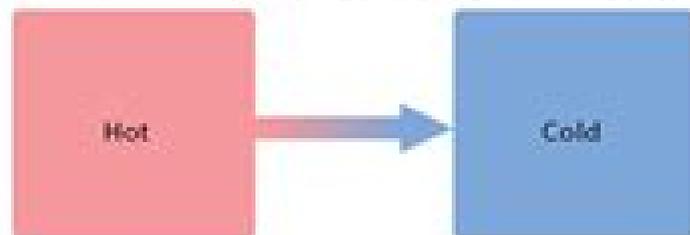
The change in internal energy ( $\Delta U$ ) of a system equals to the heat added to the system minus the work done



## Second Law

The entropy ( $S$ ) of any natural and spontaneous process either increases or remains constant

Example: Heat flow from a hot body to a cold body

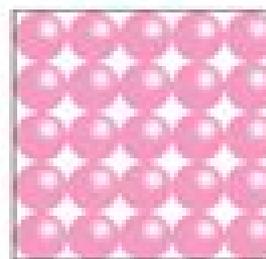


$\Delta S = 0$  For reversible process

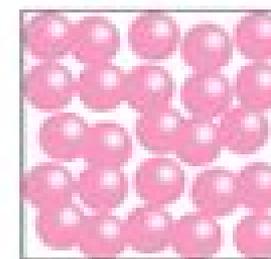
$\Delta S > 0$  For irreversible process

## Third Law

Entropy ( $S$ ) of a pure crystal is zero as the temperature ( $T$ ) approaches absolute zero



$T = 0$   
 $S = 0$



$T > 0$   
 $S > 0$

# Thermodynamics Of One Component Systems

**L Darling-Hammond**

A decorative graphic element consisting of a light blue horizontal bar with a rounded right end, and a red-to-white gradient semi-circle overlapping the right side of the bar.

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